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(54) Title: BEVERAGE STABILIZATION			
(57) Abstract			
<p>The present invention relates to a method for the removal of polyphenols and proteins simultaneously from a beverage by contacting the beverage with an ion exchanger that is capable of adsorbing both types of substances. The characteristic feature of the ion exchanger to be used is that it is a water-insoluble porous hydrophilic matrix to which ion exchanging groups are covalently bound.</p>			

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Beverage stabilization

Technical field

The present invention relates to beverage stabilization, more precisely to a method for stabilization of beverages by 5 removing haze-forming substances using an ion exchanger.

Background of the invention

The quality of beverages is measured by different parameters, such as flavour stability, biological purity and 10 physico-chemical stability, wherein the latter is one of the most important for beer. The physico-chemical or colloidal stability describes the occurrence of non-biological haze in bottled beverages, such as beer. This haze is mainly caused by polyphenols and proteins, which are able to react to 15 larger molecules via hydrogen bridges. The haze-forming proteins are believed to have Mw within the range of 30-60 kDa, although the range may differ depending on source. When removing proteins with Mw above 120 kDa, the head retention of the beer decreases. Members of the polyphenol group are, 20 inter alia, tannins and anthocyanogens which measured separately can be used as indicators for beer stabilization. To increase the beer's stability it is usual to remove partly the polyphenols, the proteins or both by using various agents and methods. The normally desired shelflife for stabilized 25 beer is about 6 months, with variations for different countries and/or kind of beer.

Stabilization and clarification of beverages are two terms that sometimes are used interchangeable and sometimes as distinct concepts. In a strict sense clarification refers to 30 the removal of haze and particulate matter that are at hand in a given beverage while stabilization refers to the removal of potentially haze-forming substances in order to render haze formation more difficult. In the context of the present invention the two terms shall be interpreted in a 35 strict sense.

The amount of haze-forming substances and their tendency to form haze depends on several factors. See the experimental part. Each beer, for instance, is unique in composition depending on the brewery's selection of process variables,

quality of hop and barley etc. This means that an acceptable level of stability/stabilization as measured by commonly accepted tests may vary between type of beer and/or brewery.

In connection with the invention fixed limits for

5 stabilization are therefore hard to be set. As a general guideline it can be said that stabilization has occurred when a value for a test measuring both haze-forming proteins and polyphenols has been changed at least 10% towards stabilization as a consequence of employing the invention.

10 This means that a similar change or even lower change may apply to tests measuring proteins and polyphenols separately. The goal with stabilization is not to remove all haze forming proteins and/or polyphenols, because this might easily also affect the character of a specific beverage.

15

Background publications

The problems with haze-causing substances in beverages have been known for several years and a number of solutions for removal thereof have been suggested.

20 The most common way to remove polyphenols from beverages is to use polyvinyl pyrrolidone, PVPP. Before addition to the beverage, PVPP has to be mixed with water to form a slurry. PVPP is added to the storage tank or dosed into the beer stream before the beer filter and filtered out with other 25 haze particles. PVPP is available in two qualities: single use and re-usable differing in particle size.

From SU 1 451 159 a process is known for beer stabilization using an ion exchange sorbent for removing polyphenols. In this process, the beer is warmed to 65-75°C, and a strongly 30 basic macro crosslinked sorbent (anion exchanger) based on copolymers of styrene and 4% divinyl compound containing functional quaternary trimethylamine groups is added. This hydrophobic sorbent is added to beer in amounts ensuring the removal of 25-30% of the polyphenols, stirred, left to stand 35 2-3 mir., and the beer is decanted. The sorbent has a polyphenol sorption capacity of 18-19 mg/g, can be reused for at least ten times, and is regenerated with 5 parts per volume water at 45-50°C.

From Rep. Res. Lab. Kirin Brew.Co. (1972), No. 15, 17-24 an anion exchanger (Dowex 1 x 4 resin) is used for fractionation of polyphenols in beer. First the polyphenols are extracted from beer, such as by ethyl acetate, then this extract is 5 subjected to anion exchange chromatography in the purpose of fractionating the polyphenols into several groups for further studies. Thus, the anion exchanger is not used for stabilization of beverages.

In Europe, one of the most common ways to remove proteins 10 from beer is by using silica which has to be prepared to a slurry with water. The silica gel is added to the storage tank or is constantly dosed into the beer stream before the beer filter. The silica gel is filtered out with other haze particles and is wasted after use.

15 Another widely used method to remove proteins is by using tannic acid. After preparing a solution of tannin and water, the tannin is added to the storage tank or is constantly dosed into the beer stream before the beer filter. The tannin is filtered out with other haze particles and is wasted after 20 use.

Also, there exist alternative methods for removing proteins from beverages containing haze-forming substances, such as by using proteolytic enzymes or bentonite.

The fact that the widely used silica gel and tannins are 25 not re-usable, makes them substantial contributors to environmental pollution.

During the priority year the Swedish Patent Office has issued an International Search Report citing the following publications:

- 30 a. US-A-4,100,149 that deals with the removal of proteins from beer and other beverages in order to accomplish, for instance, a clarified beer. The adsorbent used is built up of polymer coated inorganic particles. The polymer carries ion exchange groups. The experimental 35 part focuses on coated silica particles, and it is not clear if the effect achieved is due to the silica, polymer coat or charged groups.

- b. EP-A-166,238 deals with neutralizing the bacteriostatic activity of polyphenols in fruit juices by addition of an agent that may or may not have ion exchange groups.
- c. Hughes, Food Technology in New Zealand, 10(30) (1985)
- 5 suggests that cellulosic ion exchangers could be used to stabilize beer because they are known to adsorb proteins.
- d. US-A-4,288,462 suggests that a filter element charged with anionic colloidal silica can be used for removing haze and haze-forming proteinaceous substances in beverages, such as beer.
- 10 e. US-A-3,623,955 concerns removal a certain enzyme from beer.
- f. US-A-3,940,498 suggests to use an acid treated synthetic magnesium silicate for the simultaneous removal of undesirable proteins and polyphenols from beverages such as beer.
- 15 g. WPI acc. no. 89-212564/29 is the same as SU 1,451,159 which has been discussed above.

20 Some additional publications are:

- h. WPI acc. no 81-81725D (= DD-A-150,078) suggests the use of particulate hydrophobic material exhibiting groups capable of H bridge formation, ion exchange, chemisorption or chelation for the removal of turbidity-forming materials from drinks, including beer.
- i. WPI acc. no. 81-15897D (GB-A-2,056,485) suggests using positively charged particles for removing substances causing cloudiness in beverages (wine, beer, fruit juices etc). The charge has been introduced by treating the particles with a cationic polyamide-polyamine epichlorohydrin synthetic material. When contacted with the beverage the particles provoke precipitation that later can be removed by filtration.
- j. WPI acc. no. 77-09751Y (= GB-A-1,499,849) suggests to remove haze precursors in beverages by the use of cation exchangers based on hydrophobic matrixes.

Summary of the invention

The present invention provides a method for the removal of polyphenols and proteins simultaneously from a beverage by contacting the beverage with an ion exchanger that is capable 5 of adsorbing both types of substances. The characteristic feature of the ion exchanger to be used is that it is a water-insoluble porous hydrophilic matrix to which ion exchanging groups are covalently bound.

Physically the matrix may be in the form of a fixed bed 10 consisting of packed porous beads/particles or a porous monolith or a membrane (continuous matrixes). The shape of the beads/particles may be spherical or irregular. In the alternative the matrix may be in the form of a fluidized bed that may be unmixed (classified, stabilised, expanded) in 15 order to allow chromatography, or completely mixed as used in batch-wise procedures.

The matrix may be built up of a polymeric network exposing hydrophilic groups, such as hydroxy groups and/or amide groups, on the surface that is to contact the beverage during 20 the inventive stabilization method, i.e. both on outer surfaces and on pore surfaces. Suitable polymers are mostly organic and of biological origin (biopolymers), although also fully synthetic polymers are also contemplated. Examples of useful biopolymers are polysaccharide gels made from dextran 25 (Sephadex®, Pharmacia Biotech AB, Uppsala Sweden), agarose (Sephadose®, Pharmacia Biotech AB, Uppsala Sweden), starch, cellulose (Sephacel®, Pharmacia Biotech AB, Uppsala Sweden) etc that have been substituted with the appropriate ion exchanging groups and possibly also cross-linked. Appropriate 30 examples of synthetic polymers are polymers of hydroxyalkyl acrylates or methacrylates, hydroxy alkyl vinyl ethers, acryl or methacryl amides that optionally are N-substituted etc. The above-mentioned biopolymers and synthetic polymers have a pronounced hydrophilic character because they carry hydroxy 35 and/or amide groups along their polymer chain. Also purely hydrophobic polymers, such as polystyrenes including styrene-divinyl benzene copolymers, may be used. In the latter case it becomes imperative that the pore surfaces of the matrix have been hydrophilized, for instance by being coated

(physical adsorption or grafted) with a substance that provides the appropriate hydrophilicity, for instance the above mentioned hydroxy-group containing polymers or a low molecular weight hydroxy group containing compound (SOURCE™,

5 Pharmacia Biotech AB, Uppsala, Sweden).

The matrix, in particular in case it is in beaded form, may contain inorganic material, although as suggested above the main constituents should be of organic origin, i.e. > 50% by weight (saturated with beer or water).

10 It is important that the porosity of the gel is sufficiently high to allow for penetration of the beverage destabilizing proteins and polyphenols. Accordingly the gel should be permeable to haze forming proteins and polyphenols and hence permeable to globular proteins below 10^7 , often 15 below 5×10^6 Dalton. The ion exchange capacity is typically within the range 0.05 - 0.50 mmol per mL packed bed.

The ion exchanging groups may be cation exchanging or anion exchanging. Examples of cation exchanging groups are carboxy ($-COO^-$), sulphonic acid ($-SO_3^-$), phosphonic acid groups etc.

20 Examples of anion exchanging groups are quarternary, tertiary, secondary and primary amino groups ($-N^+(R_1, R_2, R_3)$). The free valence indicates a covalent link to the matrix and is typically through an organic spacer structure, for instance pure alkylene or hydroxy alkylene. $R_{1,3}$ is typically 25 hydrogen or lower alkyl (C_{1-6}) that may be substituted with one or more hydroxy groups. Among the quarternary amino groups may, in particularly, be mentioned $-N^+(CH_3)_3$, that when linked to the matrix via the spacer $-CH_2CHOHCH_2-$ is called a Q-group. In the experimental part SP Sepharose® and CM 30 Sepharose® are used. CM (carboxymethyl) in CM Sepharose® stands for $-OCH_2COO^-$ that is substituting an OH group directly linked to the base matrix (agarose). SP (sulfopropyl) stands for $-(CH_2)_3SO_3^-$ that is substituting an OH group of the base matrix (agarose) via a linker - 35 $OCH_2CHOHCH_2O^-$.

The steps of the inventive method comprises a. contacting the matrix with the beverage to be stabilized with one of the above-mentioned ion exchangers under conditions permitting adsorption of the destabilizing proteins and phenols, b.

recovering the beverage from the ion-exchanger. Typically one also includes an optional third step c. meaning regeneration of the ion exchanger by contacting it with a solution for regeneration, for instance containing sodium hydroxide and 5 sodium chloride and water only. As indicated above the contacting may be carried out by allowing the beverage to pass through a vessel (e.g. a column) containing the ion exchanger in any of the above-mentioned forms. Typically the regeneration solution is allowed to pass through in a 10 direction that is opposite to the direction of the beverage. The process may be continuous or in a batch-wise form. After the beverage has been in contact with the ion exchanger there often is no imperative need of a further filtration step. Compare the experimental part. Typically for beer, the 15 temperature, at least during the contact with the ion-exchanger, is above the freezing point for beer and below +10°C such as below +5°C. Most breweries today work at about 0°C, which for practical reasons is preferred also in the inventive method. Also other steps conventionally used within 20 the technical field of beverage stabilization may be included.

Best Mode

The best mode known at the priority date of the inventive 25 method comprises the regeneration step as described above. The ion exchanger used is an anion exchanger of the Q-type (Q-Sepharose®) in form of macro beads (mean particle size of 200 µm) packed in column through which the beverage is allowed to flow. Q-Sepharose® is permeable to globular 30 proteins of 4×10^6 Dalton and has an ion exchange capacity of 0.18-0.25 mmol per mL packed bed. The temperature is around 0°C. See also Example 4.

Detailed description of the invention

35 The invention will now be described more closely below in association with some non-limiting Examples. The Examples are all related to beer but it is to be understood that the method according to the invention is equally applicable on

other solutions, for example beverages other than beer, containing haze-forming substances.

EXPERIMENTAL PART

5 Materials and methods

In the Experiments, the following analysis were carried out to describe the stability of beer.

TABLE A

Analysis	Determination	Litterature/Source
Ammonium sulphate precipitation	Haze-causing protein	MEBAK*, 1993, p. 164-165
Alcohol-chill-test (ACT)	Haze-causing polyphenols and proteins. Beer stability..	MEBAK*, 1993, p.160-162
Polyphenols, total	Polyphenol content	MEBAK*, 1993, p.169-170
Anthocyanogens	Anthocyanogen content	MEBAK*, 1993, p.171-172
Accelerated Ageing test, 0/40°C	Beer stability	MEBAK*, 1993, p.157-158

10

* MEBAK is an abbreviation of Mitteleuropäische Brautechnologische Analysenkommission.

1. Protein sensitive tests:

15 1.1. Ammonium sulphate precipitation. When adding a saturated ammonium sulphate solution to beer, a haze of precipitated proteins is formed. The more ammonium sulphate solution is needed to cause this haze, the more the beer is stabilized.

20 1.2. "Esbach-test". High molecular weight proteins will be precipitated with "Esbach-reagent" (picric acid - citric acid solution). The addition of the reagent will cause a haze which can be measured photometrically.

2. Polyphenols

2.1. Polyphenols total. All polyphenols, preferably those with vicinal hydroxyl groups, are measured. The

5 polyphenols react in caustic solution with iron ions to coloured iron complexes, which can be measured photometrically.

2.2. Anthocyanogens are phenolic substances which will be turned into red-coloured anthocyanidines by the

10 treatment of hydrochloric acid.

3. Tests for the determination of beer stability.

3.1. Alcohol-chill-test. When chilling beer, a reversible haze is formed, caused by precipitated polyphenol-

15 protein complexes. The addition of alcohol decreases the solubility of these complexes and accelerates the precipitation.

3.2. Accelerated ageing test. Beer is stored at 0°C and 40°C or 60°C until haze of 2 EBC units can be recognized. The 20 haze is caused by the precipitation of polyphenol-protein complexes.

Beer stability is a complex feature and depends on several variables, including protein and/or polyphenol content. The 25 Alcohol-chill-test and the Accelerated-ageing test show a linear relation to stabilization. The tests under item 1 and 2 above do not show a linear realtion.

The following ion exchangers were tested for use in the method according to the invention.

TABLE B

Ion exchanger	Functional group	Exclusion limit (Daltons)	Mean bead size (μm)
<u>Anion exchangers:</u>			
Q Sepharose® Fast Flow	quaternary ammonium	4×10^6	90
Q Sepharose® Big Beads	-"-	-"-	200
SOURCE™ 30Q	-"-	$\leq 10^7$	30
DEAE Sephadex®	diethyl amino ethyl	1×10^6	100
<u>Cation exchangers:</u>			
SP Sepharose® Fast Flow	sulpho propyl	4×10^6	90
SP Sepharose® Big Beads	-"-	4×10^6	90
CM Sepharose® Fast Flow	carboxy methyl	4×10^6	90

5 All the Experiments were run at a temperature of about 0°C and without any pretreatment of the filtered beer.

Experiment 1

For Experiments 1 and 2, a chromatography column with an inner diameter of 50 mm (Pharmacia XK 50/1.0) was packed with 60 ml Q Sepharose® Big Beads. Then several liquids were pumped through the column according to the following C.I.P. (cleaning in place) program:

- 800 ml water, 10 min.
- 15 - 1000 ml 1 M NaOH, 60 min.
- 500 ml water, 10 min.
- 500 ml 2M NaCl, 30 min.

- 500 ml water, 10 min.

This program was also carried out after each run with beer for regeneration of the columns. 30 l filtered and unstabilized beer was pumped through the packed column with a flow rate of 6 l/h. This beer was collected, analyzed and the data compared with untreated beer. The results are given in Table I below.

TABLE I

Treatment of beer	ml NH ₄ SO ₂ / 100 ml beer	ACT EBC units	Poly- phenols mg/l	Anthocyano- genes mg/l	Accel Ageing test, days 40°C
untreated	10	10.8	205	49	1.5
Q Sepharose®	15	4.4	164	35	18

10

* EBC is an abbreviation of European Brewery Convention. Reference for EBC unit is a mixture of hydrazine sulphate and hexamethylene tetramine (formazine). 1 absolute unit = 9 000 units Helm = 15 500 EBC units.

15

Experiment 2

50 l filtered and unstabilized beer was pumped through the packed column with a flow rate of 7.8 l/h. This beer was collected, analyzed and the data compared with untreated beer. The results are given in Table II below.

TABLE II

treatment of beer	ml NH ₄ SO ₂ per 100ml beer	ACT EBC units	Poly- phenols mg/l	Antho- cyano- genes mg/l	Accel. Ageing test, days 40°C
untreated	7	10.2	201	54	1.2
Q Sepharose®	11	8	188	45	8

As appears from Table I and II, beer treated with Q Sepharose® had less polyphenols and anthocyanogens than the untreated beer. The protein-sensitive ammonium sulphate precipitation showed a reduction of protein in the treated beer. Furthermore, the results of the alcohol-chill-test and the ageing test showed that the colloidal stability of beer treated with Q Sepharose® was clearly better than the corresponding values for the untreated beer.

10

Experiment 3: Comparison of different ion exchangers

In this Experiment, for each ion exchanger, a column with an inner diameter of 10 mm was packed with 3 ml ion exchanger, washed and equilibrated with 100 ml buffer (ethanol 4.5 % v/v, adjusted to pH 4.5 by citric acid). Then 500 ml of filtered, unstabilized beer was pumped through each packed column.

Silica gel (FK700 from Degussa, Germany) was used for comparative purposes.

20 The thus treated beer was collected, analyzed and the datas are given below in Table III and IV, respectively.

TABLE III

Sample	Haze (EBC) after adding 15 ml NH ₄ SO ₄ /100 ml (EBC units)
untreated beer	10
silica gel	1,1
Q Sepharose® Fast Flow	4,2
Q Sepharose® Big Beads	4,8
SP Sepharose® Fast Flow	6
SP Sepharose® Big Beads	6
CM Sepharose® Fast Flow	10

25 As appears from the Table, besides silica gel, Q Sepharose® led to the highest reduction of haze-causing proteins in beer. SP Sepharose® also had some stabilizing effect while CM Sepharose® had no effect at all compared with untreated beer.

TABLE IV

Sample	Haze after adding 15 ml NH ₄ SO ₄ /100 ml	Alcohol-chill test (EBC units)
untreated	10.0	25.0
silica gel	1.7	4.8
SOURCE™ 30Q	4.2	18
DEAE Sephadex®	4.7	5.6
Q Sepharose® Big Beads	4.7	4.5
Q Sepharose® Fast Flow	5.2	3.0

It appears that SOURCE™ 30Q had the best protein adsorbing properties but the result of the alcohol-chill test with this product was poor. Q Sepharose® Big Beads and DEAE Sephadex® had nearly comparable stabilizing properties and the good results of the alcohol-chill test indicated a good stability of the treated beer by the additional adsorption of polyphenols. Compared to DEAE Sephadex®, Q Sepharose® has a higher chemical stability, which is an important advantage for the cleaning and sanitation process in the brewing industry.

The preferred ion exchanger according to the invention is Q Sepharose® Big Beads in respect of beer stabilizing properties, permeability and chemical stability.

Experiment 4: Large scale beer stabilization

This experiment was carried out in a brewery under practice conditions. In this brewery the beer is usually stabilized by using 15 g PVPP/hl and 30 g silica gel/hl.

For experiment 4 a chromatography column with an inner diameter of 30 cm (Pharmacia BPG 300) was packed with 9 liter of Q Sepharose® Big Beads. The bed height of Q Sepharose® in the packed column was 13 cm. Then several liquids were pumped through the column (C.I.P. program):

- 200 liter water, 10 min.
- 30 liter 2 M NaCl, 20 min.
- 40 liter water, 15 min.

- 30 liter 1 M Na OH, 120 min.
- 40 liter water, 15 min.
- 30 liter 1 M NaCl, 20 min.
- 40 liter water, 15 min.

5 This program was also carried out after each run with beer.

145 hl filtered an unstabilized beer was pumped through the packed column with a flow rate of 10 hl/h. This beer was collected and analyzed. The datas were compared with untreated beer and with beer stabilized as usual from the 10 normal production (15 g PVPP/hl and 30 g silica gel/hl). The results are given in the table V below.

Table V

treatment of beer	ml NH ₄ SO ₄ / 100 ml beer	ACT, EBC units	Anthocyanogenes mg/l	Accelerated Ageing test, days at 40°C
untreated	11	15.1	65	1.4
Q Sepharose®	13	11.2	48	7.8
normal (15 g PVPP/hl) (30 g silica/hl)	19	10.9	52	8

15 As appears from the table, beer treated with Q Sepharose® had less anthocyanogens than the untreated beer. The protein-sensitive ammonium sulphate precipitation showed a reduction of protein in the treated beer. The results of the alcohol-chill test and the ageing test showed that the colloidal 20 stability of beer treated with Q Sepharose® was much better than the untreated beer. Compared with the normal treatment for stabilization of beer, Q Sepharose® has adsorbed slightly less protein but more anthocyanogens. The stabilizing effect, described with the alcohol-chill test and the ageing test, 25 shows comparable results. This means that beer treated with Q Sepharose® had the same colloidal stability as the PVPP and silica stabilized beer, whereas the stabilization with Q Sepharose® offers the following advantages:

1. Q Sepharose® combines the effect of protein and polyphenol adsorbing materials. Therefore only one material and step is necessary instead of two for a combined beverage stabilization.
- 5 2. Q Sepharose® is reusable, whereas a reusable protein adsorbing material and a combined reusable protein and polyphenol adsorbing material does not exist.
3. Because Q Sepharose® is reusable for more than 170 cycles, it causes much less environmental pollution
- 10 4. The beverage stabilization with Q Sepharose® is easy to handle and can be automated.
5. The stabilization process is separated from beverage filtration. This allows a simple and combined beverage stabilization independent from current and future
- 15 6. Compared with i.e. enzymes used for beverage stabilization, Q Sepharose® is non-soluble.

20 A comparison of the flow rate with a common used kieselguhr filter is given in table VI below.

Table VI

	specific flow rate, hl/m ² filter area·h
kieselguhr filter	1-2
Q Sepharose® packed column	141

25 As appears from the table, columns packed with Q Sepharose® allow very high flow rates.

C L A I M S

1. A method of stabilizing beverages containing haze causing substances, comprising the following steps:

- 5 a) contacting the beverage with a water-insoluble porous hydrophilic matrix to which ion exchanging groups are covalently bond;
- b) recovering the beverage from the matrix; and,
 optionally,
- c) regeneration of the matrix.

10

2. A method according to claim 1, wherein the matrix is a fixed bed consisting of packed porous beads/particles, a porous monolith or a membrane.

15 3. A method according to claims 1 or 2, wherein the matrix comprises a polymeric network exposing on its surfaces hydrophilic groups, preferably hydroxy groups.

4. A method according to claim 3, wherein the matrix is made
20 of a hydrophilic polymer, preferably selected from polysaccharides, polymers of hydroxyalkyl acrylates or methacrylates, polymers of hydroxyalkyl vinyl ethers, and polymers of acryl or methacrylamides that optionally are N-substituted.

25

5. A method according to any one of the above claims,
 wherein the matrix is permeable to globular proteins
 below 10^7 Daltons.

30 6. A method according to any one of the above claims,
 wherein the ion exchanging groups are anion exchanging groups.

7. A method according to claim 6, wherein the anion
35 exchanging groups are quaternary ammonium groups, in particular a Q-group.

8. A method according to claims 6 or 7, wherein the matrix is based on agarose to which Q groups has been linked through $-OCH_2CHOHCH_2O-$.

5 9. A method according to any one of the claims 1-5, wherein the ion exchanging groups are cation exchanging groups.

10. A method according to any one of the above claims, wherein the method is run at a temperature below 10°C,
10 preferably below 5°C, and, for stabilization of beer, preferably above the freezing point for beer.

11. A method according to any one of the above claims, wherein the method is run continuously.

15

12. Use of a water-insoluble porous hydrophilic matrix to which ion exchanging groups are covalently bond for stabilization of beverages.

INTERNATIONAL SEARCH REPORT

International application No.

PCT/SE 97/00548

A. CLASSIFICATION OF SUBJECT MATTER		
IPC6: C12H 1/04 According to International Patent Classification (IPC) or to both national classification and IPC		
B. FIELDS SEARCHED		
Minimum documentation searched (classification system followed by classification symbols)		
IPC6: C12H, C12G, A23L, A23C Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched SE,DK,FI,NO classes as above		
Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)		
MEDLINE, BIOSIS, EMBASE, CA, WPI, CLAIMS, JAPIO, FOODSCIENCE, FOODLINE		
C. DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 4100149 A (FRANCOIS MEILLER ET AL), 11 July 1978 (11.07.78), column 1, line 1 - line 14; column 3, line 4 - line 15, claims and example 12 --	1-12
A	EP 0166238 A1 (KIRIN BEER KABUSHIKI KAISHA), 2 January 1986 (02.01.86), see claims --	1-12
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<input checked="" type="checkbox"/> Further documents are listed in the continuation of Box C.		<input checked="" type="checkbox"/> See patent family annex.
* Special categories of cited documents: "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed		
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Date of the actual completion of the international search		Date of mailing of the international search report
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INTERNATIONAL SEARCH REPORT

International application No.

PCT/SE 97/00548

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